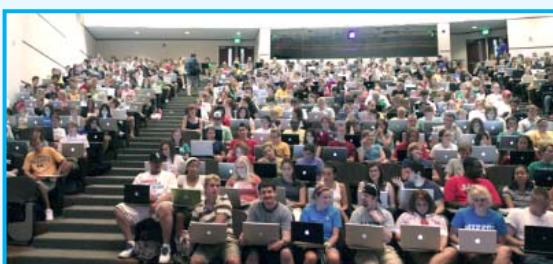


No More Lecturing! What John Falconer Has Done for Student Learning

Milo Koretsky



Concept Warehouse

Cyber-enabled infrastructure for conceptual questions

- Create a community of Learning focused on concept-based instruction
- Lower the activation barrier to promote implementation of concept-based instruction and active learning



The AIChE Concept Warehouse

AIChE[®]
Education Division
**CONCEPT
WAREHOUSE**

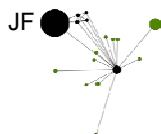


At Launch,
2012 ChE Summer School

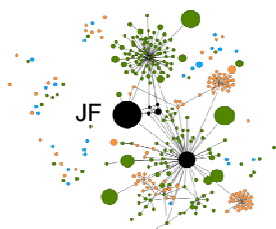


Concept Warehouse: Propagation of ConceptTests

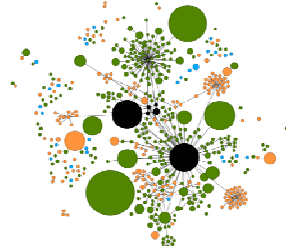
(a) Launch



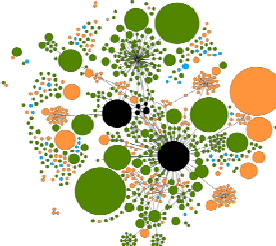
(b) 1 Year After Launch



(c) 2 Years After Launch



(d) 3 Years After Launch



Gilbuena, Smith, and Koretsky, *Educational Technology Research and Development* (2016)



Example Question: What students see when you assign online

Question text

An ideal gas flows steadily through the piping system and valve shown below. The inlet pressure and temperature are P_1 and T_1 , and the pressure drops through the valve to a lower value, P_2 .

Plus Figure

Assuming the valve is well insulated and inlet and outlet pipes connected to the valve are the same diameter, what is the relationship of the outlet temperature T_2 to the inlet temperature T_1 ?



- ☐ $T_2 > T_1$ because work is done on the gas as it is compressed through the valve opening
- ☐ $T_2 < T_1$ because temperature must decrease if pressure decreases since the volume and number of moles both stay the same
- ☐ $T_2 = T_1$ because rapid expansion of an ideal gas does not affect temperature
- ☐ Can't answer unless the type of gas flowing is specified

Written Reflection (optional)

Please explain your answer in the box below

Confidence (optional)

Please rate how confident you are with your answer:

substantially unsure moderately confident neutral confident

Submit

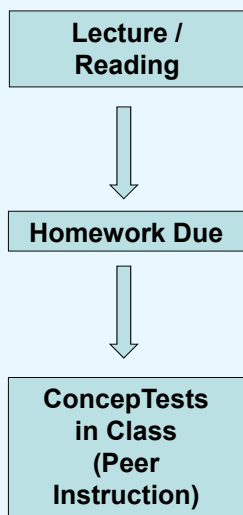
Multiple Choice Answer Selections

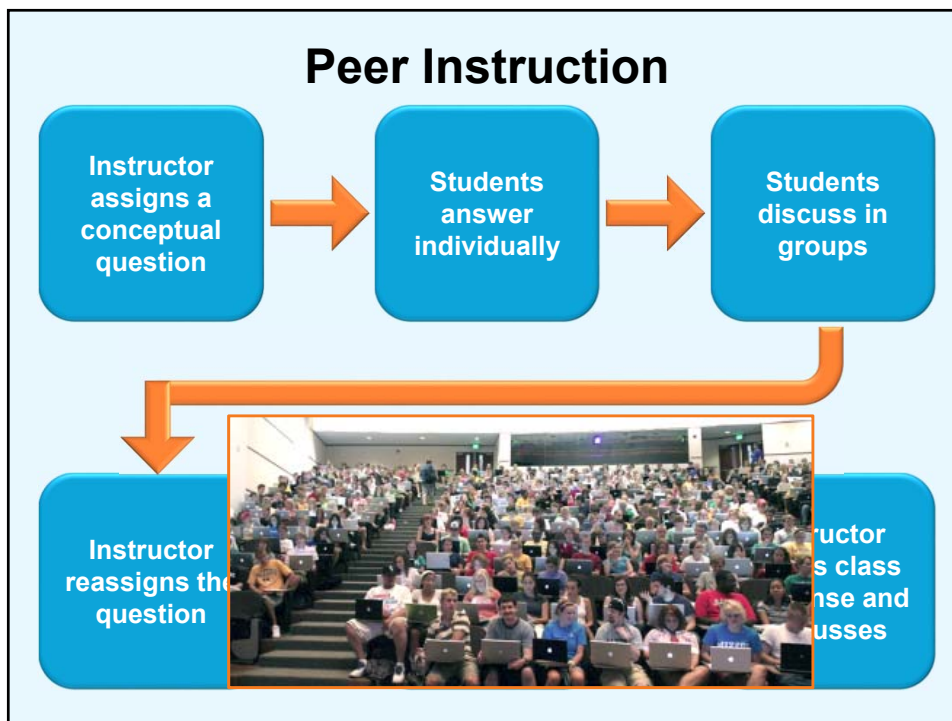


Instructional Design



Instructional Sequence





Falconer's Advantages – Peer Instruction

- Students teach/learn from fellow students
 - encourage cooperation
 - more engaged class
 - students hear alternate explanations
- Students determine how well they understand
- Instructor gets feedback from everyone
- Students motivated to be prepared; attendance higher
- Students articulate reasoning

*Methods and Tools to Help Students Learn Core Chemical Engineering Concepts
2012 ASEE Summer School*

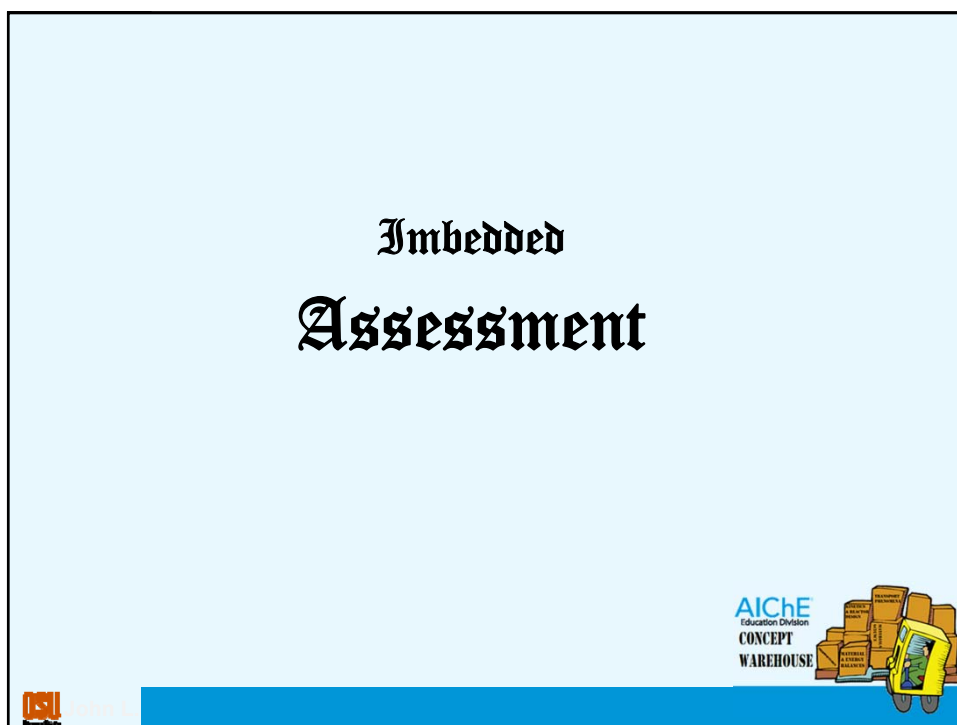
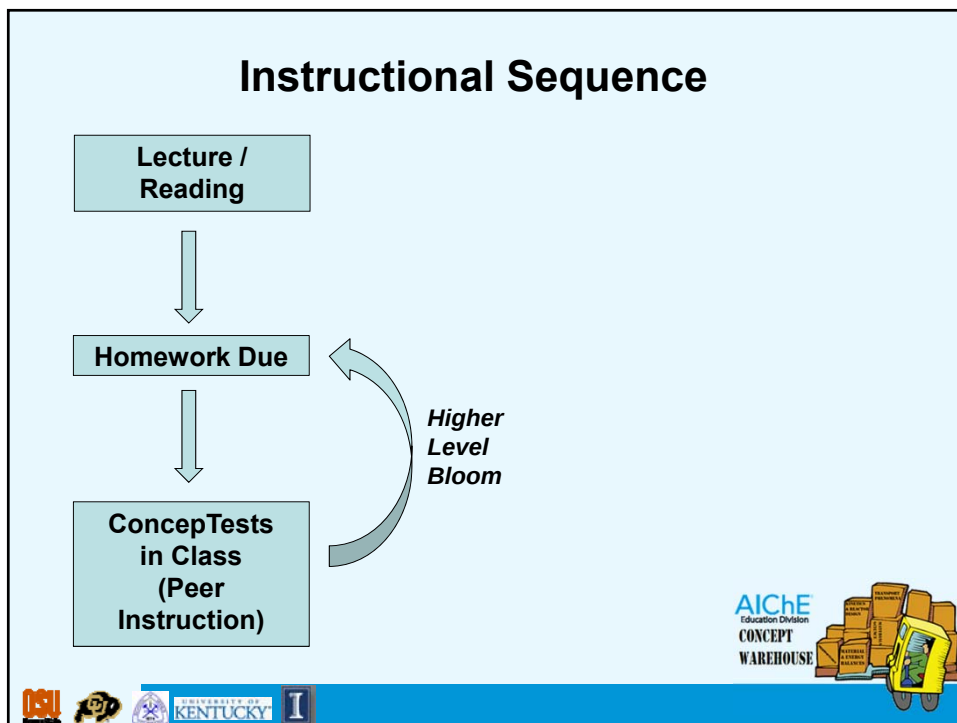
Milo Koretsky

John Falconer

Ron Miller

David Silverstein
UNIVERSITY OF
KENTUCKY

Marina Miletic
THE UNIVERSITY OF
NEW MEXICO



Instructional Sequence (Assessment)

Lecture /
Reading



Homework Due



ConcepTests
in Class
(Peer
Instruction)



Homework – traditional (Year 1)

①

55
60

Brewery

m_1 10.0% mol ethanol
90.0% mol H_2O

$m_2 = 5.00 \times 10^4 \text{ kg/h}$
 $e_2 \text{ kg E/h}$ $w_2 \text{ kg W/h}$

- distillation

water

$m_4 = 1/10 \text{ of } m_2$
 $e_4 \text{ kg E/h}$ $w_4 \text{ kg W/h}$

15% wt ethanol
55% wt H_2O

$m_3 = 1/10 \text{ of } m_2$

$e = \text{mass flow rate ethanol}$
 $w = \text{mass flow rate water}$
 $m = \text{mass flow rate total}$

$m_2 = m_3 + m_4$
 $e_2 = e_3 + e_4$
 $w_2 = w_3 + w_4$

average molecular weight
 $m_1 = \frac{.900 \text{ kmol } H_2O}{\text{kmol}} \cdot \frac{18.016 \text{ kg}}{\text{kmol } H_2O} + \frac{.100 \text{ kmol ethanol}}{\text{kmol}} \times \frac{46.07 \text{ kg}}{\text{kmol EtH}}$

avg mol weight
 $m_1 = 20.82 \text{ kg/kmol}$

Homework – electronic (Year 2)

On a warm summer day in Oslo, the temperature is 21°C and the relative humidity is 70%.

What is the absolute humidity?

Number

kg / kg Dry Air

What is the humid volume of the air?

Number

m³ / kg Dry Air

What is the specific enthalpy of the air, taking H₂O (L, 0°C, 1 atm) as a reference state?

Number

kJ / kg Dry Air

Instructional Sequence

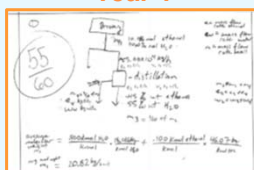
Lecture /
Reading

Homework Due

ConcepTests
in Class
(Peer
Instruction)



Year 1



Year 2

On a warm summer day in Oslo, the temperature is 21°C and the relative humidity is 70%.

What is the absolute humidity?

Number

kg / kg Dry Air

What is the humid volume of the air?

Number

m³ / kg Dry Air

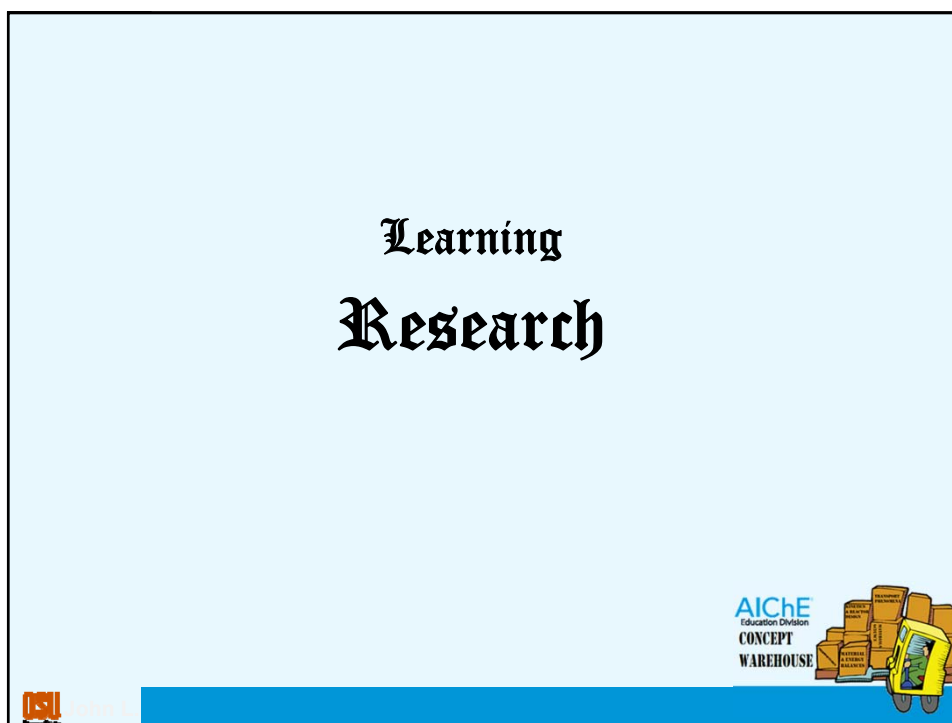
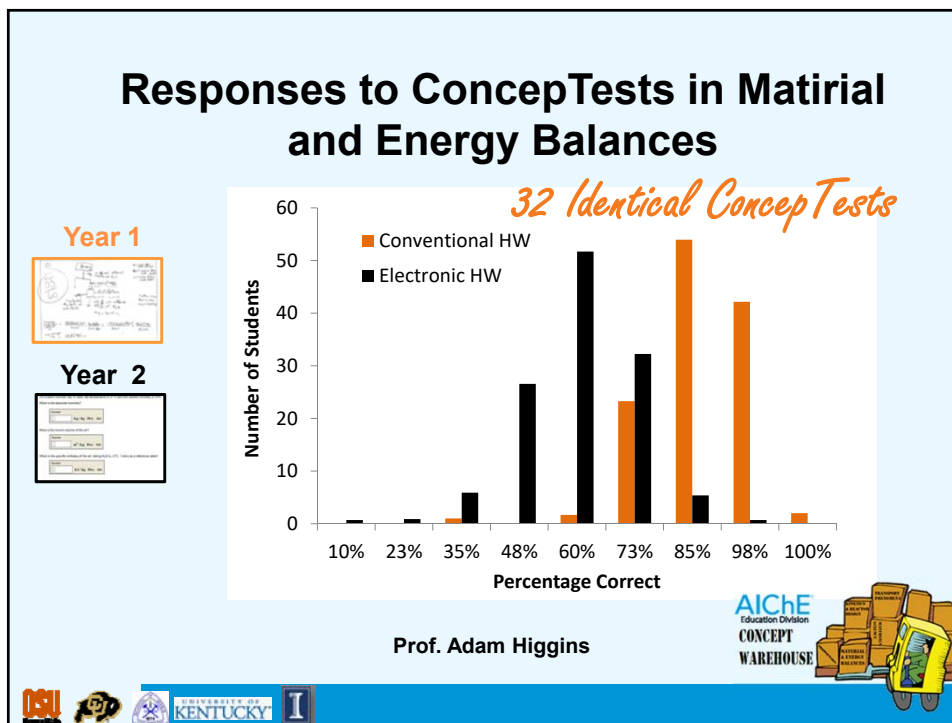
What is the specific enthalpy of the air, taking H₂O (L, 0°C, 1 atm) as a reference state?

Number

kJ / kg Dry Air



Collect Data
Here



Example Question: What students see when you assign online

Question text

Plus Figure

Written Reflection (optional)

Confidence (optional)

HOME
QUESTIONS
PROFILE

An ideal gas flows steadily through the piping system and valve shown below. The inlet pressure and temperature are P_1 and T_1 and the pressure drops through the valve to a lower value, P_2 .

Assuming the valve is well insulated and inlet and outlet pipes connected to the valve are the same diameter, what is the relationship of the outlet temperature T_2 to the inlet temperature T_1 ?

☐ $T_2 > T_1$ because work is done on the gas as it is compressed through the valve opening
☐ $T_2 < T_1$ because temperature must decrease if pressure decreases since the volume and number of moles both stay the same
☐ $T_2 = T_1$ because rapid expansion of an ideal gas does not affect temperature
☐ Can't answer unless the type of gas flowing is specified

Please explain your answer in the box below:

Please rate how confident you are with your answer:

substantially
unsure
moderately
unsure
neutral
moderately
confident
substantially
confident

Multiple Choice Answer Selections

Influence of Writing Justifications (Tr) vs. not (Com)

- Split, Crossover Design
 - Same lecture
 - Two large recitation sections
 - One section writes explanation = Treatment (Tr)
 - One section does not = Comparison (Com)
 - Alternate sections every 2 – 3 weeks
 - Compare answers to 39 ConcepTests across two cohorts
 - Chi-squared test, $\alpha = 0.05$
 - For Com group, qualitatively code reasoning (1= poor; 4 = well reasoned)



Influence of Writing Justifications (Tr) vs. not (Com)

Effect	n questions Expected
Tr > Com	~1
Tr = Com	~37
Tr < Com	~1

If the treatment
had no effect, we
would expect:

39 Questions
 $\alpha = 0.05$

Koretsky, Brooks, and
Higgins, *International
Journal of Science
Education* (2016)



Influence of Writing Justifications (Tr) vs. not (Com)

Effect	n questions	Treatment	Comparison	Treatment	Comparison
		<%>	<%>	<n students>	<n students>
Tr > Com	15	65%	50%	66	68
Tr = Com	19	58%	57%	66	67
Tr < Com	5	56%	69%	69	69

39 Questions
 $\alpha = 0.05$

Koretsky, Brooks, and
Higgins, *International
Journal of Science
Education* (2016)



How Many Students Can Justify Their Answer?

Effect		Easy	Moderate	Difficult
		$\bar{x} > 75\%$	$75\% > \bar{x} > 50\%$	$\bar{x} < 50\%$
Tr > Com	(% correct)	(85%)	(65%)	(23%)
	N _{correct}	72	39	18
	N _{code=4}	53	21	7
Tr < Com	(% correct)	NA	(59%)	(22%)
	N _{correct}		47	16
	N _{code=4}		11	1



Influence of Writing Justifications (Tr) vs. not (Com)

Effect	n questions	Treatment	Comparison	Treatment	Comparison
		<%>	<%>	<n students>	<n students>
Tr > Com	15	65%	50%	66	68
Tr = Com	19	58%	57%	66	67
Tr < Com	5	56%	69%	69	69

39 Questions
 $\alpha = 0.05$

Expanding Piston
(Falconer)

Balloon Rising
(Someone else)

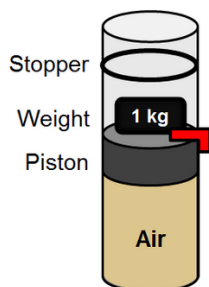
Isomorphic
Questions

Koretsky, Brooks, and
Higgins, *Journal of
Engineering Education*,
forthcoming



Expanding Piston (Falconer)

Air at high pressure and ambient temperature is contained in a perfectly insulated piston-cylinder device. If the locks holding the piston in place are removed, the piston moves upwards to a stopper. The temperature of the air _____.



- ☐ increases
- ☐ remains the same
- ☐ decreases
- ☐ need more information

Please explain your answer in the box below.

Balloon Rising

A perfectly insulated balloon filled with an ideal gas rises into the sky. As the balloon rises, the external pressure decreases, causing the balloon to expand. What happens to the temperature of the ideal gas inside the balloon?

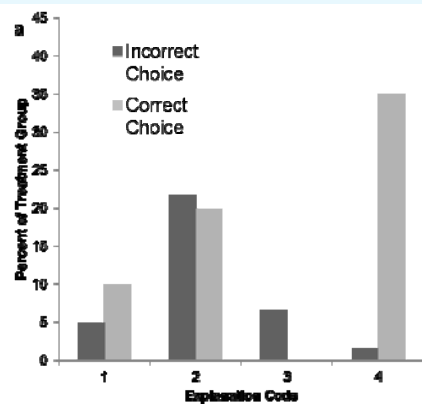
- ☐ increases
- ☐ decreases
- ☐ remains the same
- ☐ need more information

Please explain your answer in the box below.

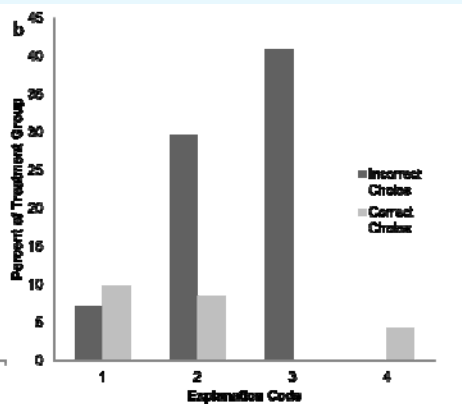
Submit

Looking at the Explanations

Expanding Piston (Falconer)



Balloon Rising



Koretsky et al., *Journal of Engineering Education* (2016)



An ideal gas at high pressure and ambient (room) temperature is contained in a perfectly insulated piston-cylinder device. The locks holding the piston in place are removed, and the piston moves upwards against ambient atmosphere to a stopper. What happens to the temperature of the ideal gas inside the piston-cylinder device?

- ☐ increases
- ☐ decreases
- ☐ remains the same
- ☐ need more information

Please explain your answer in the box below.

Submit

Expanding Piston (Modified)

Koretsky et al., *Journal of Engineering Education* (2016)



Collaborators

Student Researchers

Debra Gilbuena, PostDoc
 Bill Brooks, PostDoc
 Christina Smith, PhD
 Kritsa Chinandon, PhD
 Alec Bowen, HBS
 Rachel White, BS
 Daniel Reid, BS
 Matt Boggess, BS
 Cole Morgan, BS
 Matt Kirsch, BS

Faculty

John Falconer, University of Colorado
 Adam Higgins, Oregon State University
 Steve Krause, Arizona State University
 Carl Lira, Michigan State University
 Marina Miletic, Miletic Educational Consulting
 Ron Miller, Colorado School of Mines
 Mike Prince, Bucknell University
 Brian Self, Cal Poly SLO
 David Silverstein, University of Kentucky
 Margot Vigeant, Bucknell University

- **Beta Testers**
- **The developing community who has contributed to and used the Concept Warehouse**



Acknowledgements

- **National Science Foundation**
 - DUE - 1023099, 1022957, 1022875, 1022785
 - DUE - 1245482,
 - DUE - 1225456 (Krause Lead)
 - DUE - 1225221 (Vigeant Lead).
- **LL Stewart Scholar Program**
- **Technology Resource Program**



*Any opinions, findings, and conclusions or recommendations expressed in this material are those of the author(s) and do not necessarily reflect the views of the National Science Foundation.



AICHE
Education Division
**CONCEPT
WAREHOUSE**



<http://cw.edudiv.org>

